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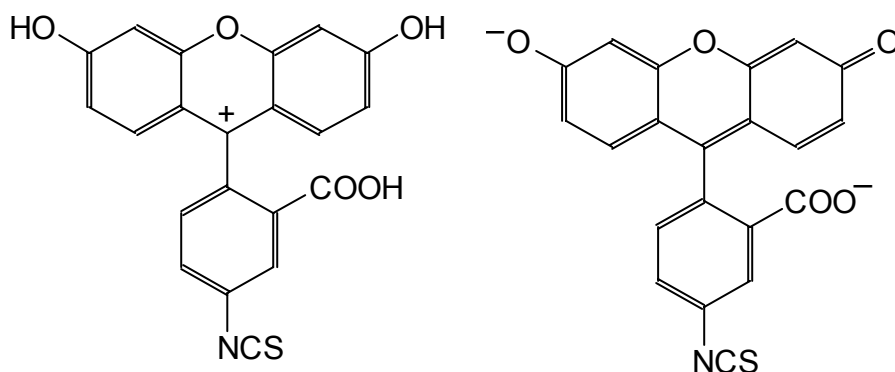
**ACID-BASE BEHAVIOR OF FLUORESC EIN ISOTHIOCYANATE IN AQUEOUS MEDIA AND IN MICELLAR SURFACTANT SOLUTIONS<sup>1</sup>**

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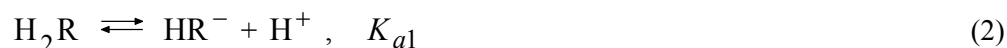
The experimental data characterizing the acid-base behavior of fluorescein isothiocyanate (FLITC) in water and in micellar solutions of a cationic surfactant N-cetylpyridinium chloride are presented. Absorption spectra are tabulated within a wide pH region at ionic strength 0.05 M (298 K). The  $pK_a$  values of stepwise dissociation of FLITC ( $H_3R^+ \rightleftharpoons H_2R \rightleftharpoons HR^- \rightleftharpoons R^{2-}$ ) are calculated.

**1. Introduction**

Fluorescein isothiocyanate (FLITC) belongs to hydroxyxanthene dyes, which are widely used owing to their unique optical properties. Protolytic (acid-base and tautomeric) equilibria of fluorescein dyes have been intensively investigated for decades.<sup>1-21</sup> However, FLITC still was not studied in details, notwithstanding its wide utilization in fluorosensors and for other purposes.<sup>22-24</sup> Therefore, it is of interest to clarify the behavior of FLITC not only in water, but also in organized solutions. Thus, in the present communication we report the data characterizing the protolytic equilibria of FITC in water and in micellar solutions of a cationic surfactant, N-cetyl pyridinium chloride, CPC. The cationic and dianionic structures of the dye,  $H_3R^+$  and  $R^{2-}$ , respectively, are shown below.



Preliminary Vis-spectroscopic studies revealed, in accord with our expectations, that for this dye, ionization in aqueous solutions occurs stepwise:



The spectrophotometric method was used to determine dissociation constants,  $K_a$ .

**2. Experimental**

Fluorescein-5-isothiocyanate ('Isomer I'),  $C_{21}H_{11}NO_5S$ , 389.38, Spiro[isobenzofuran-1(3H),9'-(9H)xanthen]-3-one, 3',6'-dihydroxy-5-isothiocyanato- from Sigma (approx. 98%, HPLC) was kindly provided by Dr. Sergei A. Eremin, Department of Chemical Enzymology, Faculty of Chemistry, M.V. Lomonosov Moscow State University. CPC (Merk, 98%) was used as received.

To create the required pH values, analytical-grade hydrochloric acid, sodium hydroxide, acetic and phosphoric acids and sodium chloride were used. The ionic strength ( $I$ , M) of the solutions was, as a

<sup>1</sup> The manuscript is archived in: <http://preprint.chemweb.com/physchem/0309005>.

rule, constant: in the buffer solutions appropriate amounts of NaCl stock solutions were added to maintain the total  $I = 0.05$  M, except solutions with higher HCl concentrations. In CPC solutions, the ‘true’ ionic strength was somewhat higher due to the presence of the surfactant. However, such small deviations from  $I = \text{const}$  display practically no influence on the equilibrium state.<sup>17</sup> The spectra of the  $\text{H}_3\text{R}^+$  species were obtained in 2 and 3 M HCl solutions. The standard aq. sodium hydroxide solution was prepared using  $\text{CO}_2$ -free water and kept protected from the atmosphere. The pH values of solutions were checked by means of potentiometry by using cells with a liquid junction (1 M KCl), glass electrode ESL-63-07, and a silver / silver chloride reference electrode, according to the compensation scheme on a potentiometer P 363/3 and pH meter-millivoltmeter pH-121. All the solutions were prepared and pH measurements performed at  $298.0 \pm 0.1$  K. Standard buffer solutions (pH 1.68, 3.56, 4.01, 6.86, and 9.18) were used for cell calibration. The experimental uncertainty in the measured pH value did not exceed 0.02 pH unit (standard deviation). The spectra of the dye solutions were measured by using SP-46 apparatus (of USSR origin). The working dye concentrations were as a rule ca.  $2 \times 10^{-6}$  M, and in several cases somewhat higher. All the spectrophotometric experiments were performed at  $298 \pm 1$  K.

### 3. Results

#### 3.1 Determination of $\text{p}K_a$ values

##### 3.1.1 Aqueous solutions

The studies were carried out with a series of 38 solutions within the wavelength range  $\lambda = 400$ -550 nm and the acidity range from 3 M HCl to pH 12; dye concentration was  $1.84 \times 10^{-6}$  M, optical path length 5 cm. Most of the data are presented in Table 1. At fixed  $\lambda$  the dependence of absorbance,  $A$ , vs. pH can be described by eq. (4):<sup>16-21</sup>

$$A = \frac{A_{\text{H}_3\text{R}^+} (a_{\text{H}^+})^3 + A_{\text{H}_2\text{R}} (a_{\text{H}^+})^2 K_{a0} + A_{\text{HR}^-} a_{\text{H}^+} K_{a0} K_{a1} + A_{\text{R}^{2-}} K_{a0} K_{a1} K_{a2}}{(a_{\text{H}^+})^3 + (a_{\text{H}^+})^2 K_{a0} + a_{\text{H}^+} K_{a0} K_{a1} + K_{a0} K_{a1} K_{a2}} \quad (4)$$

( $a_{\text{H}^+} = 10^{-\text{pH}}$ ). In this case only the  $A_{\text{H}_3\text{R}^+}$  and  $A_{\text{R}^{2-}}$  values can be measured directly at the appropriate acidity. The band of cation  $\text{H}_3\text{R}^+$  has  $\lambda_{\text{max}} = 441$  nm, with molar absorptivity  $E_{\text{H}_3\text{R}^+} = 54.56 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ , while the band of dianion  $\text{R}^{2-}$  is characterized by  $\lambda_{\text{max}} = 488$  nm and  $E_{\text{R}^{2-}}$  value, equal to  $86.5 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ .

As a first approximation, the fluorescein dissociation constants,  $K_{a0}$ ,  $K_{a1}$ , and  $K_{a2}$ , were obtained by using the procedure of successive iterations, described earlier.<sup>16</sup> The spectrum at pH 3.78 was taken as the  $\text{H}_2\text{R}$  spectrum, and the  $\text{p}K_{a0}$  value was calculated within the pH range 1.31-2.47 ( $I = 0.05$  M), at  $\lambda = 435, 440,$  and  $445$  nm:  $\text{p}K_{a0} = 2.04 \pm 0.04$ . Hence, at  $\text{pH} \geq 4$  the concentration of cationic species is negligible. Then, the  $\text{p}K_{a1}$  and  $\text{p}K_{a2}$  values were calculated jointly with the  $A_{\text{HR}^-}$  values at pH 3.96-6.68 ( $I = 0.05$  M) and 20 wavelengths, from  $\lambda = 420$  nm to 515 nm, by using the CLINP program.<sup>18,19,20,21</sup> The next step was the refinement of  $A_{\text{HR}^-}$  values, by using eq. 5:

$$A_{\text{HR}^-} = A + (A - A_{\text{H}_2\text{R}}) a_{\text{H}^+} K_{a1}^{-1} + (A - A_{\text{R}^{2-}}) (a_{\text{H}^+})^{-1} K_{a2} \quad (5)$$

where  $A$  is absorption at the current pH value. The  $A_{\text{H}_2\text{R}}$  values were, in turn, calculated by using eq. 6:

$$A_{\text{H}_2\text{R}} = A + (A - A_{\text{H}_3\text{R}^+}) a_{\text{H}^+} K_{a0}^{-1} + (A - A_{\text{HR}^-}) (a_{\text{H}^+})^{-1} K_{a1} \quad (6)$$

**Table 1.** Absorbances ( $A$ ) of FLITC in aqueous solutions; 298 K,  $I = 0.05$  M (at pH > 1.3)

pH	$\lambda$ , nm																			
	420	425	430	435	440	445	450	455	460	465	470	475	480	485	490	495	500	505	510	515
H <sub>3</sub> R <sup>+</sup>	0.297	0.349	0.407	0.466	0.502	0.469	0.375	0.244	0.137	0.072	0.037	0.021	0.013	0.011	0.009	0.008	0.007	0.007	0.007	0.007
1.13	0.281	0.333	0.389	0.440	0.469	0.435	0.349	0.227	0.129	0.069	0.036	0.021	0.020	0.011	0.011	0.009	0.009	0.007	0.005	0.005
1.31	0.277	0.325	0.380	0.428	0.451	0.417	0.328	0.220	0.119	0.067	0.034	0.022	0.015	0.012	0.010	0.007	0.006	0.004	0.003	0.003
1.49	0.274	0.323	0.379	0.423	0.438	0.400	0.312	0.206	0.117	0.067	0.036	0.026	0.021	0.018	0.015	0.012	0.009	0.007	0.007	0.004
1.65	0.273	0.317	0.369	0.407	0.412	0.374	0.292	0.192	0.112	0.067	0.042	0.032	0.026	0.022	0.021	0.017	0.014	0.011	0.010	0.009
1.82	0.267	0.312	0.359	0.393	0.395	0.354	0.272	0.184	0.109	0.069	0.048	0.039	0.032	0.031	0.027	0.021	0.017	0.014	0.012	0.010
1.99	0.266	0.307	0.349	0.377	0.374	0.328	0.253	0.173	0.109	0.074	0.054	0.048	0.044	0.037	0.035	0.028	0.023	0.022	0.016	0.015
2.23	0.250	0.287	0.325	0.350	0.342	0.294	0.228	0.155	0.101	0.070	0.055	0.049	0.045	0.042	0.032	0.027	0.020	0.016	0.014	0.011
2.47	0.242	0.277	0.308	0.330	0.319	0.276	0.212	0.147	0.101	0.076	0.062	0.058	0.052	0.047	0.040	0.034	0.025	0.019	0.014	0.013
3.78	0.185	0.208	0.223	0.231	0.224	0.213	0.194	0.174	0.158	0.152	0.149	0.148	0.140	0.119	0.093	0.064	0.042	0.028	0.019	0.015
3.96	0.182	0.203	0.220	0.225	0.227	0.220	0.209	0.192	0.178	0.171	0.172	0.171	0.159	0.138	0.106	0.073	0.047	0.030	0.020	0.015
4.00	0.177	0.197	0.213	0.227	0.235	0.238	0.237	0.224	0.213	0.208	0.209	0.210	0.199	0.166	0.125	0.084	0.052	0.032	0.021	0.014
4.15	0.178	0.198	0.213	0.224	0.231	0.231	0.223	0.213	0.199	0.195	0.195	0.195	0.179	0.153	0.116	0.078	0.051	0.031	0.019	0.015
4.36	0.175	0.199	0.213	0.225	0.238	0.244	0.244	0.235	0.223	0.220	0.222	0.220	0.204	0.175	0.133	0.087	0.056	0.035	0.021	0.016
4.60	0.180	0.201	0.218	0.234	0.250	0.265	0.268	0.262	0.252	0.249	0.252	0.250	0.234	0.194	0.148	0.097	0.061	0.039	0.025	0.018
4.82	0.179	0.197	0.215	0.233	0.256	0.273	0.282	0.278	0.270	0.269	0.273	0.270	0.254	0.213	0.161	0.105	0.064	0.039	0.023	0.015
5.00	0.174	0.194	0.211	0.231	0.256	0.277	0.290	0.287	0.278	0.279	0.286	0.286	0.265	0.223	0.169	0.112	0.069	0.040	0.024	0.015
5.21	0.171	0.189	0.206	0.229	0.257	0.282	0.294	0.294	0.290	0.288	0.297	0.298	0.278	0.242	0.182	0.120	0.074	0.042	0.024	0.013
5.48	0.167	0.185	0.203	0.225	0.253	0.282	0.298	0.300	0.295	0.298	0.308	0.312	0.298	0.264	0.209	0.143	0.089	0.048	0.028	0.018
5.60	0.162	0.179	0.199	0.220	0.249	0.277	0.294	0.298	0.297	0.298	0.308	0.317	0.307	0.273	0.218	0.155	0.097	0.055	0.030	0.017
5.80	0.159	0.177	0.195	0.218	0.246	0.278	0.297	0.304	0.303	0.308	0.321	0.335	0.333	0.307	0.297	0.180	0.116	0.064	0.035	0.020
5.97	0.156	0.175	0.192	0.215	0.246	0.276	0.297	0.305	0.307	0.314	0.330	0.347	0.354	0.331	0.277	0.204	0.129	0.068	0.037	0.021
6.00	0.160	0.176	0.194	0.216	0.244	0.277	0.298	0.312	0.316	0.325	0.347	0.369	0.380	0.367	0.316	0.238	0.153	0.085	0.045	0.028
6.22	0.146	0.162	0.178	0.201	0.231	0.261	0.288	0.303	0.312	0.326	0.354	0.385	0.415	0.418	0.375	0.286	0.190	0.108	0.054	0.032
6.58	0.135	0.148	0.166	0.187	0.215	0.246	0.278	0.301	0.316	0.334	0.372	0.418	0.469	0.492	0.458	0.355	0.235	0.133	0.072	0.038
6.68	0.126	0.138	0.153	0.173	0.199	0.234	0.265	0.294	0.316	0.339	0.379	0.440	0.504	0.548	0.523	0.417	0.270	0.150	0.078	0.041
R <sup>2-</sup>	0.073	0.082	0.093	0.105	0.128	0.163	0.213	0.266	0.312	0.354	0.413	0.520	0.664	0.777	0.796	0.684	0.466	0.258	0.128	0.060

Note. The H<sub>3</sub>R<sup>+</sup> spectrum is measured in solutions with  $I > 0.05$  M: C (HCl) = 1 and 2 M.

The R<sup>2-</sup> spectrum is measured at pH = 9.58; 10.00; 11.79; 11.95.

**Table 2.** Absorbances ( $A$ ) of FLITC in CPC solutions; 298 K,  $I = 0.05$  M (at pH > 1.3)

pH	$\lambda$ , nm																		
	435	440	445	450	455	460	465	470	475	480	485	490	495	500	505	510	515	520	530
H <sub>3</sub> R <sup>+</sup>	0.509	0.550	0.541	0.452	0.325	0.211	0.135	0.087	0.066	0.055	0.049	0.047	0.046	0.046	0.046	0.046	0.046	0.046	0.045
1.15	0.369	0.395	0.374	0.297	0.197	0.117	0.063	0.036	0.024	0.019	0.016	0.015	0.012	0.011	0.010	0.008	0.009	0.009	0.009
1.33	0.312	0.330	0.308	0.242	0.162	0.099	0.057	0.032	0.024	0.017	0.017	0.015	0.013	0.012	0.009	0.009	0.007	0.007	0.007
1.59	0.206	0.213	0.196	0.153	0.102	0.061	0.037	0.024	0.019	0.017	0.015	0.013	0.011	0.007	0.007	0.007	0.004	0.004	0.004
1.87	0.115	0.117	0.107	0.084	0.059	0.039	0.027	0.021	0.017	0.014	0.014	0.015	0.012	0.012	0.008	0.006	0.004	0.004	0.003
3.78	0.064	0.066	0.072	0.079	0.085	0.085	0.079	0.074	0.073	0.075	0.077	0.072	0.058	0.044	0.027	0.019	0.011	0.007	0.005
3.89	0.089	0.093	0.104	0.115	0.121	0.121	0.112	0.108	0.108	0.112	0.112	0.104	0.085	0.063	0.041	0.023	0.018	0.012	0.007
4.10	0.121	0.129	0.143	0.159	0.172	0.169	0.158	0.149	0.148	0.154	0.156	0.147	0.122	0.093	0.061	0.040	0.025	0.017	0.010
4.15	0.114	0.120	0.132	0.148	0.156	0.155	0.144	0.138	0.140	0.141	0.144	0.137	0.115	0.085	0.058	0.038	0.026	0.018	0.011
4.34	0.129	0.138	0.153	0.171	0.184	0.182	0.171	0.160	0.161	0.166	0.172	0.162	0.137	0.101	0.068	0.042	0.027	0.016	0.010
4.57	0.150	0.159	0.177	0.202	0.216	0.213	0.203	0.192	0.192	0.199	0.206	0.195	0.166	0.128	0.090	0.059	0.035	0.022	0.010
4.78	0.180	0.192	0.211	0.235	0.258	0.250	0.237	0.227	0.227	0.239	0.243	0.234	0.203	0.162	0.123	0.085	0.054	0.039	0.022
5.00	0.190	0.203	0.224	0.258	0.270	0.270	0.258	0.246	0.250	0.261	0.270	0.264	0.235	0.196	0.150	0.109	0.069	0.044	0.021
5.20	0.192	0.206	0.227	0.256	0.274	0.276	0.265	0.257	0.262	0.276	0.286	0.282	0.261	0.231	0.186	0.135	0.088	0.053	0.023
5.52	0.186	0.199	0.222	0.249	0.270	0.276	0.269	0.266	0.274	0.290	0.305	0.314	0.308	0.286	0.243	0.178	0.117	0.068	0.025
5.65	0.169	0.180	0.199	0.227	0.249	0.261	0.264	0.265	0.277	0.295	0.321	0.345	0.366	0.367	0.334	0.257	0.166	0.094	0.027
5.90	0.148	0.159	0.177	0.204	0.229	0.247	0.260	0.273	0.287	0.310	0.345	0.390	0.440	0.472	0.452	0.361	0.237	0.135	0.034
6.10	0.139	0.150	0.166	0.192	0.216	0.240	0.258	0.277	0.295	0.321	0.362	0.423	0.498	0.544	0.523	0.424	0.278	0.160	0.042
6.19	0.126	0.138	0.155	0.177	0.203	0.231	0.253	0.278	0.298	0.323	0.369	0.435	0.523	0.582	0.577	0.465	0.319	0.179	0.044
6.44	0.110	0.120	0.135	0.155	0.182	0.215	0.246	0.281	0.303	0.334	0.385	0.469	0.582	0.658	0.662	0.545	0.375	0.204	0.049
6.79	0.094	0.105	0.117	0.138	0.166	0.206	0.243	0.286	0.312	0.347	0.400	0.498	0.635	0.740	0.750	0.633	0.440	0.248	0.063
R <sup>2-</sup>	0.063	0.073	0.085	0.102	0.128	0.172	0.223	0.274	0.308	0.341	0.400	0.509	0.672	0.801	0.854	0.762	0.535	0.312	0.072

Note. The H<sub>3</sub>R<sup>+</sup> spectrum is measured in solutions with  $I > 0.05$  M: C (HCl) = 1 and 2 M.

The R<sup>2-</sup> spectrum is measured at pH = 8.11; 9.00; 9.94.

Finally, the CLINP program was used once more; the data for 28 working solutions with various pH and 20 wavelengths, from  $\lambda = 420$  nm to 515 nm, were utilized in the calculations with the result, referring to  $I = 0.05$  M:  $pK_{a0} = 2.05 \pm 0.03$ ;  $pK_{a1} = 4.35 \pm 0.02$ ;  $pK_{a2} = 6.62 \pm 0.01$ .

### 3.1.2 Apparent dissociation constants in micellar CPC solutions

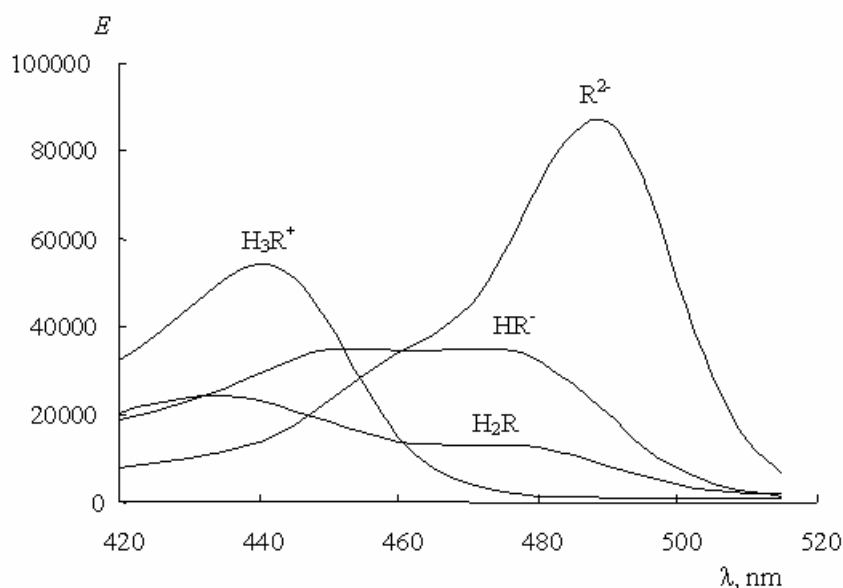
In 0.003 M CPC solutions, at bulk ionic strength  $I = 0.05$  M, the absorption of 28 working solutions of FLITC ( $2.056 \times 10^{-6}$  M) with different pH were measured at 19 wavelength,  $\lambda = 435$ -530 nm (Table 2). The determination of  $pK_{a1}^a$  and  $pK_{a2}^a$  values was made by the aforementioned procedure. The  $pK_{a0}^a$  value of FITC was obtained isolated from the other steps, at  $\lambda = 435, 440, \text{ and } 445$  nm, within the pH range 1.33–1.87 (HCl + NaCl), because in a more acidic region the ionic strength would be higher than 0.05 M. The final results are as follows:  $pK_{a0}^a = 1.36 \pm 0.18$ ;  $pK_{a1}^a = 4.31 \pm 0.01$ ;  $pK_{a2}^a = 6.00 \pm 0.01$ .

### 3.2 Molar absorptivities of ionic and molecular species

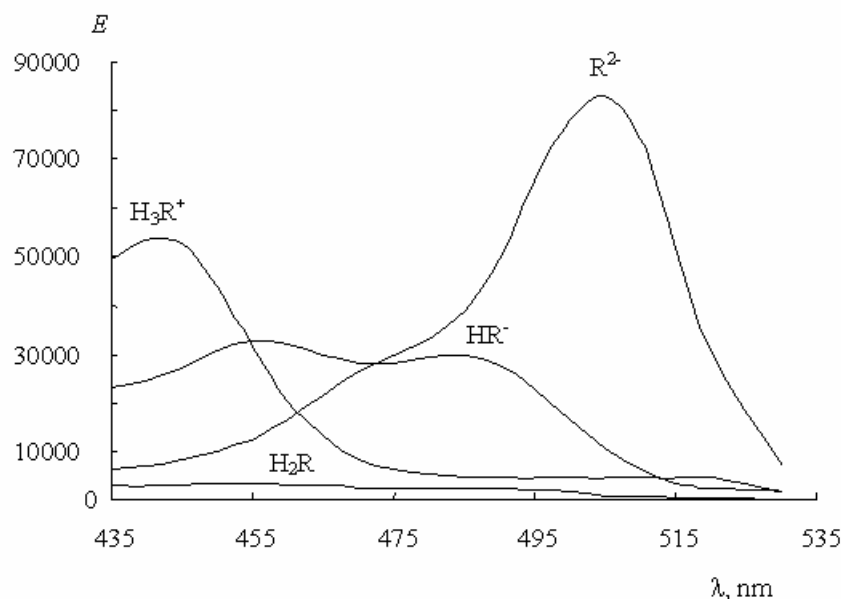
Having the  $K_{a1}$  and  $K_{a2}$  values (or  $pK_{a1}^a$  and  $pK_{a2}^a$  – in the case of micellar solutions), it was then possible to calculate the molar absorptivities of  $HR^-$  at various wavelengths, in such a way obtaining the spectra of these species (eq. 5). The interval  $pK_{a1} \leq \text{pH} \leq pK_{a2}$  was used for this purpose; such procedure can be regarded as an additional step in the procedure of  $pK_a$  and  $A_{HR^-}$  iterative calculations described above. Molar absorptivity at each wavelength was calculated as  $E_{HR^-} = A_{HR^-} c^{-1} l^{-1}$ , where  $c$  is the working dye concentration,  $l$  is the optical path length.

The refinement of  $E_{H_2R}$  values of FLITC ( $E_{H_2R} = A_{H_2R} c^{-1} l^{-1}$ ) was carried out at  $pK_{a0} \leq \text{pH} \leq pK_{a1}$ , with the help of equation (6), to avoid any influence of traces of intensely colored ions ( $H_3R^+$  and  $HR^-$ ) on the spectra of the neutral forms. The spectra were measured at different  $c$  values near to  $1 \times 10^{-5}$  M.

Thus obtained spectra of FLITC ionic and molecular species are depicted in Fig. 1, 2.



**Fig.1.** Absorption spectra of FLITC species in aqueous solutions,  $I = 0.05$  M (except the cationic spectrum).



**Fig.2.** Absorption spectra of FLITC species in CPC micellar solutions,  $I = 0.05$  M (except the cationic spectrum).

The detailed analysis of the protolytic (acid-base and tautomeric) equilibria of the dye in solutions will be published in the nearest future.

#### 4. Conclusions

1. The  $pK_a$  values of FLITC in aqueous solutions at 298 K and  $I = 0.05$  M are determined:  $pK_{a0} = 2.05 \pm 0.03$ ;  $pK_{a1} = 4.35 \pm 0.02$ ;  $pK_{a2} = 6.62 \pm 0.01$ .
2. The apparent  $pK_a^a$  values of FLITC in 0.003 M CPC solutions at 298 K and  $I = 0.05$  M are determined:  $pK_{a0}^a = 1.36 \pm 0.18$ ;  $pK_{a1}^a = 4.31 \pm 0.01$ ;  $pK_{a2}^a = 6.00 \pm 0.01$ .
3. Molar absorptivities of ionic and molecular species of the dye are evaluated in the two aforementioned solvent systems.

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Kharkov University Bulletin. 2006. №731. Chemical Series. Issue 14(37). Н. О. Мchedlov-Петросян, Ю. В. Исаенко, Н. А. Водолазкая, С. Т. Гога. Кислотно-основное поведение флуоресцеин изотиоцианата в водной среде и мицеллярных растворах ПАВ.

Представлены экспериментальные данные, характеризующие кислотно-основное поведение флуоресцеин-изотиоцианата в воде и в мицеллярных растворах катионного ПАВ *N*-цетилпиридиний хлорида. Приведены спектры поглощения в широком диапазоне pH при ионной силе 0.05 M (298 K). Рассчитаны значения ступенчатой диссоциации флуоресцеин-изотиоцианата ( $H_3R^+ \rightleftharpoons H_2R \rightleftharpoons HR^- \rightleftharpoons R^{2-}$ ).